

## Make (and Name) Your Compounds

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**Instructions: Create 20 compounds; write the formulas and names for them.**

**10 compounds can be constructed with two cut-out squares (hint: look for oxidation numbers that “match”. The other 10 must be constructed with more than two cut-out squares (hint: look for oxidation numbers that do not “match”. SEE Examples below.**

PART A Cut and Paste:

Compounds are chemically combined ions that bond together to have an overall zero charge. In order for a stable compound to form, the ions must balance so that the positive (+) ions balance with the negative (-) ions. You are going to put the positive ions and negative ions together by cutting them out from the following page, and then glue them next to each other, thus forming your compounds.

There are some ‘rules’ to follow when making your compounds:

1. Only two different types of ions will form a compound.
2. Put the positive (+) ions to the left of the negative (-) ions.
3. Be sure to leave space around your compounds so that you can name them.

PART B Write the chemical formula for your compounds:

You will write the chemical formula for your compounds below each compound you formed. To write the chemical formula, write the symbol for the first element in the compound right next to the symbol for the second element. Include subscripts and parenthesis as necessary.

PART C Write the name for your compounds:

There are some simple rules to follow when naming compounds.

1. Write the name of the cation first and the anion second.
2. Remember if the compound is binary the ending should be “ide”.
3. Include Roman Numerals in the name for the transition elements if appropriate.

Examples

